

Periodic Table Basics from General Chemistry

To get started: Know the important elements early on:

- You don't need to know all of the elements; to get started you should focus on the top of the table and elements that actually get used early on in Organic 1. Elements (e.g. transition metals) will be added later as they are needed for new reactions.
- Focus on **Carbon** being in the middle of this abbreviated table; this is important as C has a “middle” electronegativity (2.5). Elements to the left of C are less electronegative; those to the right are more electronegative. This will dictate types of bonds formed.
- Metals on the left give away electron(s) to form ionic bonds, those in the middle share, those on the right can accept electrons or share to form ionic or covalent bonds. You learned this in General Chemistry and it still applies in Organic Chemistry.

H							He
Li	Be	B	<u>C</u>	N	O	F	Ne
Na	Mg	Al	Si	P	S	Cl	Ar
K						Br	
						I	

- Before you begin Organic 1 you should know where these elements are on the table and why they are organized this way. This will save time when making decisions on how the elements will interact based on their electronegativities and whether they can share or transfer electrons. If you are looking at a periodic table during an exam to find out where an element is, you are wasting time that should be spent elsewhere.
- Notice that some elements are crossed out. We don't talk about Beryllium (it's rare) and we don't make molecules from the inert (Noble) gasses (because they are inert with a full octet). The rest of the elements in the above table will do what they did in General Chemistry so you should be on top of that as you prepare for Organic 1.